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Article
Fe-Based Metallic Glasses and Dyes in Fenton-Like Processes: Understanding Their Intrinsic Correlation

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Abstract: Fe-based metallic glasses have been demonstrated as effective heterogeneous catalysts in Fenton-like processes for dye degradation. Yet, currently corresponding studies have limitations due to the limited study object (dyes) and the correlation between metallic glasses and dye pollutants in Fenton-like processes is still not comprehensively studied. Accordingly, this work intensively investigated the thermal catalytic behavior correlations between two Fe-based metallic glasses (Fe 78 Si 9 B 13 and Fe 73.5 Si 13.5 B 9 Cu 1 Nb 3 ) and eight different dyes. Results indicated a lower activation energy in the more active metallic glass and a dependence of the activation energy of Fe-based metallic glasses in dye solutions. In addition, a high H 2 O 2 concentration led to a declined catalytic efficiency but a photo-enhanced Fenton-like process overcame this limitation at high concentration of H 2 O 2 due to the decrease of pH and enhancement of irradiation. Furthermore, the average mineralization rates of Fe 78 Si 9 B 13 and Fe 73.5 Si 13.5 B 9 Cu 1 Nb 3 have been measured to be 42.7% and 12.6%, respectively, and the correlation between decolorization and mineralization revealed that a faster decolorization in a Fenton-like process contributed to a higher mineralization rate. This work provides an intrinsic viewpoint of the correlation between Fe-based metallic glasses and dyes in Fenton-like processes and holds the promise to further promote the industrial value of metallic glasses.

Keywords: metallic glasses; dyes; Fenton-like processes; activation energy; mineralization

1. Introduction

In recent years, metallic glasses with long-range disordered atomic configurations have been extensively investigated as novel functional catalysts in environmental and energy science [1–4]. Compared to catalysts with highly ordered crystalline structures, their uniquely amorphous structures fabricated by rapid solidification processes endow metallic glasses with many fascinating catalytic properties, such as high surface active sites induced by unsaturated coordination of atomic numbers [5], fast electron transfer [6], low thermal activation energy barriers [7], self-stabilizing nature [8] and structural stability [9,10]. Particularly, low-cost Fe-based metallic glasses with their easily modified chemical composition, good glass-forming ability and environmentally friendly nature have attracted increasing interest as catalysts in renewable energy conversion. For example, Fe 40 Co 40 Se 20 supported on carbon fiber paper [11], Fe 40 Ni 40 P 20 [5,12] and Fe 40 Co 40 P 13 C 7 [13] metallic glass ribbons were directly
employed as electrocatalysts for efficient hydrogen/oxygen evolution reactions; a comparative study between metallic glasses and their crystalline counterparts also suggests that the amorphous structure easily induces a much higher active catalytic activity [12] and the functionality of amorphous structures with active sites and reduced energy barriers has been demonstrated [14]. Recently, stimulated by global environmental pressure, the attractive catalytic properties in Fe-based metallic glasses have become tremendous advantages for the treatment of environmental pollutants [15].

Advanced oxidation processes (AOPs) have been considered as effective methods to purify organic contaminants [16], which cannot be achieved by conventional chemical, physical and biological methods [17–20]. The major contribution to effective pollutant degradation from AOPs is usually attributed to the formation of reactive species with high redox potential, such as hydroxyl radical (•OH, $E^0 = 2.7$ V) [21]. Fenton/Fenton-like processes are one type of effective AOPs to generate •OH by hydrogen peroxide (H$_2$O$_2$) and iron-containing materials under acidic conditions. Given the fact that the amorphous nature facilitates Fe-based metallic glasses to have better catalytic properties than crystalline catalysts, they have been applied as a superior alternative of crystalline iron catalysts in Fenton-like processes. It has been reported that Fe$_{78}$Si$_9$Si$_{13}$ metallic glass ribbons presented 5–10 times faster of •OH production rate than other Fe-based crystalline catalysts in methylene blue and methyl orange dye degradation [22]; Fe$_{60}$Ni$_{30}$P$_{13}$C$_7$ metallic glass ribbons displayed a 2-fold enhanced catalytic efficiency by surface reactivation in a Fenton-like process [23]; Fe$_{80}$P$_{13}$C$_7$ metallic glass ribbons showed self-renewing and ultra-strong catalytic reusability in dye decomposition [24,25]. Very recently, our research group has also demonstrated that Fe$_{78}$Si$_9$B$_{13}$ metallic glass ribbons could be effective environmental catalysts not only for the aforementioned degradation of organic pollutants, but also for remediation of inorganic pollutants (arsenic and nitrate) [26], showing a high promise of industrial value.

Due to the development of economy and the increased demand of market, synthetic dyes are currently widely applied in many industries, especially in the textile industry as coloring agents [27]. According to their chemical structures and functional groups, dyes with chromophore and auxochrome groups are classified into azo dyes, triarylmethane dyes, xanthene dyes, etc. [28]. Chromophore groups determine the presence of color by absorbing light within the visible spectrum while auxochrome groups determine the color intensity [29]. Functional groups such as amino, hydroxyl, nitro and carbonyl groups can further alter the light absorbance capability of chromophore groups [30]. However, dye molecules not only contain genotoxic and carcinogenic substances but also have chemically stable structures of a recalcitrant nature [31], which will cause detrimental effects on the aquatic environment if a dye solution is discharged without effective treatment. In addition, the synthesis methods of dyes make their chemical structures more and more complicated, which also greatly increases the difficulty for degrading dyes into harmless substances. It is thus urgent to identify and exploit effective methods to alleviate the environmental pressure.

Although Fe-based metallic glasses used in Fenton-like processes have been investigated for effective dye degradation in recent years, these studies mostly focused on specific dye degradation leading to a significant limitation. In addition, the correlation between metallic glasses and dyes is not yet fully understood. Therefore, motivated by these limitations, our research group focused on two Fe-based metallic glasses (Fe$_{78}$Si$_9$B$_{13}$ and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$) in Fenton-like processes to degrade eight dyes with different classifications (i.e., azo dyes, thiazine dye, triarylmethane dyes and nitroso dye), aiming to understand the intrinsic correlation of Fe-based metallic glasses and dyes in Fenton-like processes.

2. Results and Discussion

2.1. Dye Characteristics and Properties

The ultraviolet (UV)-visible (Vis) spectra of the eight different dyes are shown in Figure 1. In order to distinguish the functional groups in their chemical structures, the UV-Vis spectra have been divided...
into the Vis and UV region (Figure 1a,b, respectively). As shown in Figure 1a, the intensities and absorbance peaks of different dyes with the same concentration (20 ppm) are apparently different, where the peak intensity of methylene blue (MB) is stronger than that of the other dyes. Cibacron brilliant yellow 3G-P (BY), cibacron brilliant red 3B-A (BR) and naphthol green B (NG) dyes have the lowest peak intensity among all eight dyes. According to the characteristic peaks of dyes in the Vis spectra, Table 1 summarizes their absorbance peak positions ($\lambda_{\text{max}}$), which generally dominate the reflection of their colors by light absorbance. In addition, all the dye structures are shown in Table 1, where their functional groups and state of charge are clearly indicated for their classification. Accordingly, methyl orange (MO), bright black BN (BB), BY and BR dyes can be classified as azo dyes due to the presence of a nitrogen double bond (–N=N–) functional group with an anionic charge and they have been widely studied in recent years [22,32–34]. MB dye with a four-carbon ring, one nitrogen and one sulfur atom is classified as a thiazine dye and it is commonly used in medications [24,35]. Nowadays many researchers tend to employ MB dye as a classical organic pollutant example, since it is relatively stable at different pH values and solution temperatures, so as to investigate effect of AOPs using different catalysts [9,36–38]. With triphenylmethane backbones, both crystal violet (CV) and malachite green (MG) dyes can be grouped as triarylmethane dyes with cationic charges and usually present intense color [39]. It should be noted that MG dye has a self-conversion between the MG molecule and MG leucocarbinol [40]. NG dye with a –N=O functional group can be classified as a nitroso dye.

![Figure 1](image_url)

**Figure 1.** (a) Visible spectra of eight dyes (i.e., MO, MB, BB, BY, BR, CV, MG and NG) as organic pollutants ([Dye]$_0$ = 20 ppm, pH = 3.0). (b) Corresponding ultraviolet spectra of the eight dyes under the same conditions.

It should be mentioned that in addition to chromophore groups, most dye molecules are attached to benzene rings or/and heterocyclic aromatic molecules, leading to additional absorbance peaks located at the UV region of UV-Vis spectra, as demonstrated in Figure 1b. All the dye absorbance spectra in the UV region contain more than one peak, indicating that the intrinsic aromatic structures are more complicated than just the chromophore groups. It is also interesting to find that some azo dyes (e.g., BR and BY) with a large molecular weight and a low peak intensity in the Vis region, however, have a high peak intensity in the UV region. In contrast, the peak intensities of triarylmethane dyes (i.e., CV and MG) in the UV region are relatively lower than the peaks in the Vis region. Although hydroxyl radicals (•OH) produced by Fenton-like processes have strong and non-selective oxidation properties towards organic pollutants [41], the intrinsically distinguishable chemical structures of dyes in fact may lead to side reactions during their degradation, and the degradation effect of dyes may be even closely related to the type of catalysts used, which is still not fully understood in the Fenton-like processes catalyzed by Fe-based metallic glasses.
Table 1. Characteristics of the dyes.

<table>
<thead>
<tr>
<th>Dye</th>
<th>Classification</th>
<th>Empirical Formula</th>
<th>Molecular Weight (g/mol)</th>
<th>$\lambda_{\text{max}}$ (nm)</th>
<th>Structure</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>Methyl Orange (MO)</td>
<td>Azo-anionic</td>
<td>C$<em>{14}$H$</em>{14}$N$_{3}$NaO$_3$S</td>
<td>327.33</td>
<td>498</td>
<td>[9,22,42]</td>
<td></td>
</tr>
<tr>
<td>Methylene Blue (MB)</td>
<td>Thiazine-cationic</td>
<td>C$<em>{16}$H$</em>{18}$ClN$_3$S</td>
<td>319.85</td>
<td>664</td>
<td>[24,43-45]</td>
<td></td>
</tr>
<tr>
<td>Bright Black BN (BB)</td>
<td>Azo-anionic</td>
<td>C$<em>{23}$H$</em>{27}$N$_3$Na$_4$O$_4$S$_4$</td>
<td>867.68</td>
<td>570</td>
<td>[23]</td>
<td></td>
</tr>
<tr>
<td>Cibacron Brilliant Yellow 3G-P (BY)</td>
<td>Azo-anionic</td>
<td>C$<em>{23}$H$</em>{29}$Cl$_3$N$_4$Na$_3$O$_4$S$_4$</td>
<td>831.02</td>
<td>404</td>
<td>[46]</td>
<td></td>
</tr>
<tr>
<td>Cibacron Brilliant Red 3B-A (BR)</td>
<td>Azo-anionic</td>
<td>C$<em>{32}$H$</em>{34}$Cl$_2$N$_4$Na$_4$O$_4$S$_4$</td>
<td>1000.25</td>
<td>517</td>
<td>[33,47]</td>
<td></td>
</tr>
<tr>
<td>Crystal Violet (CV)</td>
<td>Triarylmethane-cationic</td>
<td>C$_{28}$N$<em>3$H$</em>{30}$Cl</td>
<td>407.98</td>
<td>582</td>
<td>[16,48]</td>
<td></td>
</tr>
<tr>
<td>Malachite Green (MG)</td>
<td>Triarylmethane-cationic</td>
<td>C$<em>{23}$H$</em>{21}$ClN$_2$</td>
<td>364.91</td>
<td>618</td>
<td>[27,40,49]</td>
<td></td>
</tr>
<tr>
<td>Naphthol Green B (NG)</td>
<td>Nitroso-anionic</td>
<td>C$<em>{30}$H$</em>{15}$Fe$_3$N$_3$Na$_3$O$_3$S$_3$</td>
<td>878.45</td>
<td>714</td>
<td>[50]</td>
<td></td>
</tr>
</tbody>
</table>

2.2. Thermal Catalytic Behavior Correlation Between Metallic Glasses and Dyes

Generally, the intrinsic activation of chemical reaction by catalysts may have various behaviors and the thermal activation energy barrier ($\Delta E$) is commonly used to investigate the thermal behavior of catalysts in their reactions. Many studies have demonstrated that metallic glasses as amorphous catalysts have lower $\Delta E$ than most crystalline alloy catalysts, indicating that the amorphous nature of
the catalysts facilitates the occurrence of degradation processes \[7,51,52\]. The $\Delta E$ of metallic glasses are generally lower than 60 kJ/mol \[9\]. However, this fact does not indicate that the $\Delta E$ of the same metallic glass will be constant toward reactions in the different dyes since the reaction temperature environment may also affect degradation of dye with variation of structure so as to influence the whole process of degradation by the metallic glass, which is not fully understood. As such, Fe\textsubscript{78}Si\textsubscript{9}B\textsubscript{13} and Fe\textsubscript{73.5}Si\textsubscript{13.5}B\textsubscript{9}Cu\textsubscript{1}Nb\textsubscript{3} metallic glass ribbons have been employed to investigate their thermal catalytic behavior with the eight different dyes. In this work, the $\Delta E$ of specific metallic glass ribbons with dyes is measured from the corresponding first-order reaction rate constant ($k$) under reaction temperatures of 25, 30, 35 and 40 °C. The $\Delta E$ is obtained by the Arrhenius equation: 

$$\ln k_{\text{obs}} = -\Delta E/RT + \ln A,$$

where $k$ is first-order reaction rate constant, $R$ is the gas constant (8.314 J/(mol·K)), $T$ is the absolute temperature (K), and $A$ is a pre-exponential factor.

Figure 2a–h show the degradation performance of Fe\textsubscript{78}Si\textsubscript{9}B\textsubscript{13} metallic glass ribbons on the eight different dyes from 25 to 40 °C. It can be seen that almost all the dye decolorizations reach completion within 10 min at 25 °C, showing the outstanding catalytic performance of Fe\textsubscript{78}Si\textsubscript{9}B\textsubscript{13} metallic glass ribbons in the Fenton-like process due to their fast electron transfer ability and relatively strong surface stability, which have been intrinsically investigated in recent years \[6,9\]. However, two dyes (BY and NG) present exceptional decolorization performance, where BY reaches a 90% decolorization and the NG decolorization rate is about 80% within 10 min at the same temperature (25 °C). In addition, compared to our previous work \[50\], the NG decolorization rate in the Fenton-like process in this work is slightly lower than that in the photo-enhanced sulfate radical-based AOPs (about 95%) in previous work using the same Fe\textsubscript{78}Si\textsubscript{9}B\textsubscript{13} metallic glass ribbons. The higher decolorization rate of NG in the photo-enhanced sulfate radical-based AOPs may be attributed to the effect of the external UV-Vis irradiation effect and the different AOPs system used. Both BY and NG show negligible enhancement effect upon full completion of dye decolorization although the reaction rate is enhanced at high temperatures. This enhanced reaction rate has also been observed in the other six dyes (MO, MB, BB, BR, CV and MG) when the reaction temperature is increased. In fact, the temperature also has a similar effect as UV-Vis irradiation, which facilitates completion of faster reactions due to the supplementary activation by the external energy \[33\].

![Figure 2](image_url)

**Figure 2.** Decolorization of (a) MO, (b) MB, (c) BY, (d) BB, (e) BR, (f) CV, (g) MG and (h) NG by Fe\textsubscript{78}Si\textsubscript{9}B\textsubscript{13} metallic glass ribbons under 25–40 °C without UV-Vis irradiation (dye concentration: 20 ppm, catalyst dosage: 0.5 g/L, H\textsubscript{2}O\textsubscript{2} concentration: 1 mM, pH 3.0). Decolorization rates were measured from the $\lambda_{\text{max}}$ of corresponding UV-Vis spectra with reaction time.

On the other hand, the decolorization behaviors of Fe\textsubscript{73.5}Si\textsubscript{13.5}B\textsubscript{9}Cu\textsubscript{1}Nb\textsubscript{3} metallic glass ribbons against the dyes at 25 to 40 °C are shown in Figure 3a–h. The decolorization efficiencies of Fe\textsubscript{73.5}Si\textsubscript{13.5}B\textsubscript{9}Cu\textsubscript{1}Nb\textsubscript{3} are apparently lower than those of Fe\textsubscript{78}Si\textsubscript{9}B\textsubscript{13} seen in Figure 2 due to the lower
surface adsorption ability and stronger surface protection in Fe$_{78}$Si$_9$B$_{13}$Cu$_1$Nb$_3$ [22,47]. The inclusion of Cu and Nb will also contribute to a lower electron transfer ability [53]. In this work, most of dye decolorization by Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ reached completion within 60 min and the decolorization rates also increase with the reaction temperatures. However, the decolorization of the two triarylmethane dyes (CV and MG) seemed to be inhibited when using Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ as a catalyst (Figure 3f,g). In particular, CV dye can only be decolorized by less than 40% within 45 min and under 40 °C. The CV and MG dye molecules have the same cationic charge in solution, which makes them different from the other dye molecules with anionic charges. It should be noted that the decolorization performance of Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ on the cationic dye MB does not show any inhibitory behavior, which may be attributed to a combination of the thiazine dye structure and a molecular charge favorable to the degradation in Fenton-like processes by Fe-based metallic glasses. The inhibited decolorization in CV and MG is possibly attributed to their triarylmethane structures when Cu or/and Nb are present in metallic glasses. In fact, the molecule charge affects the dye absorption on the ribbon surface and further affects the decolorization during the processes. Considering the aforementioned catalytic degradation of both two metallic glass ribbons in Figures 2 and 3, the decolorization of dyes in Fenton-like processes can be affected by their molecular charge, and dye and catalyst structures.

Accordingly, the reaction rate constants for different reaction temperatures measured from Figures 2 and 3 have been summarized in Table 2. In addition to the Arrhenius equation, all the $\Delta E$ for different dyes can be obtained by the slope of fitting curve of $-\ln k$ as a function of 1000/RT. As shown in Figure 4, the fitting of Arrhenius curve is classified by used dyes and all the obtained $\Delta E$ values have been shown based on metallic glass ribbons. It should be noted that the fitting of NG dye using Fe$_{78}$Si$_9$B$_{13}$ only includes 25–35 °C due to the extremely fast decolorization at 40 °C. In order to make it clearer, Figure 5 summarizes all the $\Delta E$ values according to dyes and metallic glass ribbons. It is clear that $\Delta E$ of Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ is always higher than that of Fe$_{78}$Si$_9$B$_{13}$, confirming that Fe$_{78}$Si$_9$B$_{13}$ metallic glass ribbons are always easier to activate than Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ in a Fenton-like process. A range of ~10–30 kJ/mol of $\Delta E$ of difference between Fe$_{78}$Si$_9$B$_{13}$ and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ can be estimated from Figure 5. In BR dye, $\Delta E$ presents the closest value for the two metallic glasses and MB dye shows the biggest difference. However, $\Delta E$ of the same metallic glass in the degradation of different dyes varies, indicating that $\Delta E$ of a metallic glass also depends on the reaction environment (dyes in this work) and it is not always consistent for the same material. This fact also indicates that any comparison of $\Delta E$ values for different catalysts should also consider if the reaction environment is consistent.
any comparison of ΔE values for different catalysts should also consider if the reaction environment is consistent. Table 2 provides a summary of the activation energy values for different catalysts, along with the dye degradation reactions. Figure 4 illustrates the activation energy ΔE of Fe78Si9B13 (FeSiB) and Fe73.5Si13.5B9Cu1Nb3 (FeSiBCuNb) metallic glass ribbons for MO, MB, BY, BB, BR, CV, MG and NG dyes. Figure 5 shows a graph summarizing the activation energy ΔE values for these dyes. A summary of the activation energy ΔE values for different catalysts should also consider if the reaction environment is consistent.

Table 2. Reaction rate constants (k) of dye degradation by Fe78Si9B13 and Fe73.5Si13.5B9Cu1Nb3 metallic glass ribbons under different reaction temperatures. All the k values were obtained from Figures 2 and 3 with fixed catalyst dosage (0.5 g/L).

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>Reaction Temperature (°C)</th>
<th>Dye Reaction Rate Constant k (min⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>MO</td>
</tr>
<tr>
<td>Fe78Si9B13</td>
<td>25</td>
<td>0.548</td>
</tr>
<tr>
<td></td>
<td>30</td>
<td>0.711</td>
</tr>
<tr>
<td></td>
<td>35</td>
<td>1.070</td>
</tr>
<tr>
<td></td>
<td>40</td>
<td>1.410</td>
</tr>
<tr>
<td>Fe73.5Si13.5B9Cu1Nb3</td>
<td>25</td>
<td>0.061</td>
</tr>
<tr>
<td></td>
<td>30</td>
<td>0.074</td>
</tr>
<tr>
<td></td>
<td>35</td>
<td>0.120</td>
</tr>
<tr>
<td></td>
<td>40</td>
<td>0.211</td>
</tr>
</tbody>
</table>

Correlation between Metallic Glasses and H2O2 in Different pH

Generally, reaction parameters such as catalyst dosage, pH, dye concentration, and H2O2 concentration dominate the reaction efficiency in Fenton-like processes and an effective degradation of a dye solution usually involves the investigation of optimized parameters. For catalyst (metallic
Within the range of 1 mM to 50 mM, the decolorization of MO still reaches almost completion within 10 min when the H₂O₂ concentration is from 1 mM to 10 mM. However, with further increase, the decay of catalytic efficiency has been further enhanced using a H₂O₂ concentration of 50 mM. The inhibitory effect of H₂O₂ on decolorization is apparent. Further increasing the H₂O₂ concentration from 50 mM to 0.5 M results in negligible MO decolorization within 45 min. In both ribbons, a high H₂O₂ concentration (>5 mM) has an inhibitory effect on the catalytic efficiency due to scavenging effect of H₂O₂ on the *OH radicals produced, which leads to a declined catalytic activity [31].

Within the range of 1 mM to 50 mM, the decolorization of MO still reaches almost completion within 10 min. The decay of catalytic efficiency has been further enhanced using a H₂O₂ concentration of 0.5 M but the decolorization can reach more than 80% in 20 min. Comparing the variation of H₂O₂ shows a stronger inhibitory effect on the catalytic efficiency using Fe₇₈Si₉B₁₃ as a catalyst and the range of feasible concentrations is very limited (Figure 6b). Only 1 and 2 mM H₂O₂ present optimized efficiency for the full MO decolorization in 45 min. From 5 mM to 20 mM, the inhibitory effect of H₂O₂ on decolorization is apparent. Further increasing the H₂O₂ concentration from 50 mM to 0.5 M results in negligible MO decolorization within 45 min. In both ribbons, a high H₂O₂ concentration has a detrimental effect on catalytic efficiency.

The aforementioned results also indicate that the effect of H₂O₂ concentration depends on the type of catalysts and a catalyst with low catalytic ability (i.e., Fe₇₃.₅Si₁₃.₅B₉Cu₁Nb₃) is more sensitive to the variation of H₂O₂ concentration. Figure 7 has shown the corresponding reaction rate constant (k) of Fe₇₈Si₉B₁₃ and Fe₇₃.₅Si₁₃.₅B₉Cu₁Nb₃ metallic glass ribbons as a function of initial H₂O₂ concentration.
concentration ([H$_2$O$_2$])$_0$). The $k$ of Fe$_{78}$Si$_9$B$_{13}$ are almost one order of magnitude higher than that of Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$, and Fe$_{78}$Si$_9$B$_{13}$ can perform a better catalytic efficiency at a wider range of H$_2$O$_2$ concentration. However, both of them have an inhibitory catalytic effect at a H$_2$O$_2$ concentration higher than 5 mM.

Figure 7. Reaction rate constants ($k$) of Fe$_{78}$Si$_9$B$_{13}$ (FeSiB) and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ (FeSiBCuNb) metallic glass ribbons as a function of initial H$_2$O$_2$ concentration ([H$_2$O$_2$])$_0$ ranging from 1 mM to 500 mM (or 0.5 M) in the degradation of MO dye (dye concentration: 20 ppm, catalyst dosage: 0.5 g/L, pH 3.0, 25 °C).

It should be noticed that Fenton-like processes usually rely on acidic conditions to maintain a reliable catalytic performance. pH 3.0 is typically used for most iron-containing catalysts, including the Fe-based metallic glasses used in this work. Under weaker acidic conditions (e.g., pH 4.0), the catalytic efficiency will decline and it will be further inhibited under neutral conditions. Figure 8a,b show the catalytic performance of the two metallic glass ribbons under neutral conditions (pH 6.2 ± 0.1) with different H$_2$O$_2$ concentrations. Both of them exhibit unapparent decolorization of MO dye, indicating that the neutral condition is unfavorable for Fenton-like processes using metallic glass ribbons [24]. At a high concentration of H$_2$O$_2$ (0.5 M), the catalytic efficiency is also very low, with 15% and 0% of decolorization rate for Fe$_{78}$Si$_9$B$_{13}$ and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$, respectively, after 60 min (Figure 8c). However, it has been reported that Fenton-like processes with UV-Vis irradiation can highly enhance their color removal rates under neutral conditions [33]. Accordingly, Figure 9a,b suggest that a low concentration of H$_2$O$_2$ (50 mM) with UV-Vis irradiation does not effectively promote the decolorization rate but the combination of high H$_2$O$_2$ concentration and UV-Vis irradiation effectively enhances the decolorization rate. At 1.0 M H$_2$O$_2$, both metallic glass ribbons do not exhibit a big difference in MO decolorization, which reaches up to 94% and 90% for Fe$_{78}$Si$_9$B$_{13}$ and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$, respectively, within 60 min. Increasing the H$_2$O$_2$ concentration under UV-Vis irradiation leading to promoted dye degradation rate reveals that the combination of high concentration of H$_2$O$_2$ and irradiation can overcome the catalytic efficiency limitation under neutral conditions.

Further investigation indicates that in addition to a photo-enhancement effect, the improved efficiency is also attributable to the effective optimization of pH conditions with increasing H$_2$O$_2$ concentration (Figure 9c,d). Accordingly, this behavior provides a novel strategy to overcome the limitation of Fenton-like processes under neutral conditions.
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Figure 8. Effect of H2O2 concentration from 1 mM to 0.5 M using (a) Fe78Si9B13 (FeSiB) and (b) Fe73.5Si13.5B9Cu1Nb3 (FeSiBCuNb) metallic glass ribbons under neutral conditions. (c) Comparison of MO degradation by Fe78Si9B13 (FeSiB) and Fe73.5Si13.5B9Cu1Nb3 (FeSiBCuNb) at 0.5 M H2O2 and neutral condition (MO dye concentration: 20 ppm, catalyst dosage: 0.5 g/L, pH 6.2 (±0.1), 25 °C). Decolorization rates were measured from the λmax of the corresponding UV-Vis spectra with reaction time.

Figure 9. Effect of H2O2 concentrations by (a) Fe78Si9B13 (FeSiB) and (b) Fe73.5Si13.5B9Cu1Nb3 (FeSiBCuNb) metallic glass ribbons (0.5 g/L) under neutral conditions (pH 6.2 ± 0.1) and UV-Vis irradiation (7.7 μW/cm²). (c,d) pH variation along reaction time corresponding to (a,b), respectively. Decolorization rates were measured from the λmax of the corresponding UV-Vis spectra with reaction time.

2.4. Correlation between Decolorization and Mineralization

Figure 10a–h show the decolorization processes of the eight dyes using Fe78Si9B13 metallic glass ribbons during 5 min. The full wavelength UV-Vis spectra (190–800 nm) clearly indicate that most of dyes have experienced an effective destruction of their chromophore groups by the Fe78Si9B13-activated Fenton-like process [42]. It should be noted that the initial absorbance peak of NG is located at 714 nm but it has shifted to 408 nm after the Fenton-like process is activated, which is also reported by sulfate radical-based AOPs [50]. After 5 min, the intensities of the characteristic absorbance peaks of the dyes become weaker and then disappear, except for CV and MG dyes. Both CV and MG dyes have
a slower decolorization efficiency than the other dyes, which becomes more apparent when using Fe73.5Si13.5B9Cu1Nb3 metallic glass. Figure 10i–p also show the decolorization processes of the eight dyes using Fe73.5Si13.5B9Cu1Nb3 metallic glass. The decolorization efficiencies are slower and the peaks disappear after 30 min. The decay of the characteristic absorbance peaks is even slower regarding CV and MG dyes (Figure 10n,o). Considering the dye structure, it can be concluded that triarylmethane dyes are unfavorable in the Fenton-like system with Fe-based metallic glass ribbons.

Although the decolorization efficiency of dyes in the Vis spectra provide an important clue about the catalytic ability by catalysts or the catalytic efficiency of the reaction system, they only include the destruction of chromophore group in dye molecules. To fully convert organic pollutants into harmless substances (CO2, H2O, etc.), the degradation pathway is much more complicated. In Figure 10, it can be seen that the absorbance peaks at UV region also become inapparent along with peaks at the Vis region. The disappearance of UV peaks reveals that the Fenton-like process not only effectively destroys the chromophore groups in the dye structures, but only involves catalytic oxidation of the aromatic structure [55]. This behavior can be further confirmed by the TOC removal rate ([TOC]/[TOC]0), which is usually employed to evaluate the mineralization rate in AOPs [16,42].

Figure 11 shows that most of dye mineralization rates reach about 40% within 15 min when using Fe78Si9B13 metallic glass ribbons while for Fe73.5Si13.5B9Cu1Nb3 metallic glass ribbons, a much lower TOC removal rate can be seen. Both Fe78Si9B13 and Fe73.5Si13.5B9Cu1Nb3 have a relatively higher TOC removal rates close to average value using specific metallic glass ribbons. In addition, the decolorization efficiency shown in Figures 2 and 3.

\[
\text{TOC removal rate} = \frac{[\text{TOC}]_0 - [\text{TOC}]_t}{[\text{TOC}]_0} \times 100\%
\]

\[
[\text{TOC}]_0 = \text{Initial TOC concentration} \\
[\text{TOC}]_t = \text{TOC concentration at time } t
\]

Figure 10. UV-Vis spectra of (a) MO, (b) MB, (c) BY, (d) BB, (e) BR (f) CV, (g) MG and (h) NG dye degradation by Fe78Si9B13 (FeSiB) within 5 min. UV-Vis spectra of (i) MO, (j) MB, (k) BY, (l) BB, (m) BR (n) CV, (o) MG and (p) NG dye degradation by Fe73.5Si13.5B9Cu1Nb3 (FeSiBCuNb) within 30 min. (dye concentration: 20 ppm, catalyst dosage: 0.5 g/L, H2O2 concentration: 1 mM, pH 3.0).

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Although the decolorization efficiencies of dyes in the Vis spectra provide an important clue about the catalytic ability by catalysts or the catalytic efficiency of the reaction system, they only include the destruction of chromophore group in dye molecules. To fully convert organic pollutants into harmless substances (CO2, H2O, etc.), the degradation pathway is much more complicated. In Figure 10, it can be seen that the absorbance peaks at UV region also become inapparent along with peaks at the Vis region. The disappearance of UV peaks reveals that the Fenton-like process not only effectively destroys the chromophore groups in the dye structures, but only involves catalytic oxidation of the aromatic structure [55]. This behavior can be further confirmed by the TOC removal rate ([TOC]/[TOC]0), which is usually employed to evaluate the mineralization rate in AOPs [16,42].

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removal rate for MO and have a relatively lower TOC removal rate for MG. The lowest TOC removal rate occurs for CV and MG dyes using Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$. According to the TOC removal rates of the eight dyes, the average values of Fe$_{78}$Si$_9$B$_{13}$ and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ metallic glasses are calculated as 42.7% and 12.6%, respectively, indicating most of dye degradations can achieve TOC removal rates close to average value using specific metallic glass ribbons. In addition, the mineralization rate is apparently closely related to decolorization, where most of time, a fast decolorization usually leads to a higher mineralization rate, corresponding to the decolorization efficiency shown in Figures 2 and 3.

The dominant reactive species in Fe-based metallic glasses-activated Fenton-like processes have been reported to be •OH [33,56–58]. The effective decolorization and mineralization in this work suggest that part of the dye molecules involved with •OH have experienced degradation pathways including cleavage of chromophore groups, formation of intermediates, further oxidation of intermediates to become CO$_2$, H$_2$O and inorganic substances [59]. For example, the main degradation mechanism of MG dye (a triarylmethane dye) in the Fenton-like process probably involves the cleavage of the conjugated chromophore structure by attack of •OH, followed by N-demethylation reactions and opening of phenyl rings to become smaller molecules [49,60], although different reaction routes may still vary according to the specific reaction conditions; the main degradation mechanism of MO dye (an azo dye) goes through the initial destruction of the azo chromophore group (–N=N–) to form smaller intermediates with functional amine (–N(CH$_3$)$_2$) and sulfonate (–SO$_3$–) groups leading to fast decolorization, followed by the demethylation of intermediates, the hydroxylation of aromatic rings, and further ring opening [61,62]. Given the fact that a low dosage of metallic glass ribbons is used in this work, their structure and surface may have a minor direct effect on dye degradation but their catalytic activation toward H$_2$O$_2$ directly affects the decolorization and further mineralization rate of dyes due to the production rate of •OH when using different metallic glass ribbons [16,22]. The intrinsic catalytic activity of the metallic glass dominates the dye degradation efficiency in this work.

3. Materials and Methods

3.1. Materials and Chemicals

Metallic glass ribbons with nominal chemical compositions of Fe$_{78}$Si$_9$B$_{13}$ and Fe$_{73.5}$Si$_{13.5}$B$_9$Cu$_1$Nb$_3$ were supplied by Qingdao Yunlu Energy Technology Co., Ltd. (Qingdao, China), and were manufactured by the melt spinning technique. For the preparation procedures of as-spun metallic glass ribbons readers may refer to a previous report [53]. The produced ribbons generally have a
thickness of 30–40 µm and were cut into pieces of 10 × 20 mm (width and length), respectively, for the following catalytic analysis with dyes. BB, BY and BR, hydrochloric acid (HCl, 37% w/w) were purchased from Sigma-Aldrich (Sydney, Australia). MO and MB were supplied by Xilong Chemical Co., Ltd. (Shantou, China). CV and NG were supplied by Wenzhou Huaqiao Chemical Reagent Co., Ltd. (Wenzhou, China). MG was purchased from Ji’an Haomai Fine Chemical Industry Co., Ltd. (Ji’an, China). Hydrogen peroxide (H₂O₂, 30 wt%) and sodium hydroxide (NaOH, 40% w/v) were purchased from Rowe Scientific Pty Ltd. (Perth, Australia). Milli-Q water with 18.2 MΩ•cm was used for catalytic processes, dilution and cleaning.

3.2. Characterization

The internal structure of as-spun Fe₇₈Si₉B₁₃ and Fe₇₃.₅Si₁₃.₅B₉Cu₁Nb₃ metallic glass ribbons have been well characterized by X-ray diffraction (XRD), differential scanning calorimetry (DSC) and transmission electron microscopy (TEM) in our previous work [53], which demonstrated their fully amorphous nature.

3.3. Catalytic Analysis

The catalytic oxidation of dyes by Fenton-like processes was investigated with involvement of two as-spun Fe-based metallic glass ribbons and different concentrations of H₂O₂ in a thermostatic water bath. The catalytic thermal behavior of Fe-based metallic glass ribbons therefore could be studied under controlled temperatures (i.e., 25, 30, 35, 40 °C). Before each experiment, 20 ppm dye solution was freshly prepared by submitting 2 mL of pre-prepared 1000 ppm dye to dilution to 100 mL. The stirring rate was fixed at 300 rpm throughout this work. Except for the initial pH 6.2 (±0.1) of as-prepared MO dye solution, all other pH values were controlled at 3.0 as initial pH by adding diluted HCl solution. The pH values in Figure 9c,d were measured after addition of specific concentration of H₂O₂ without Fe-based metallic glasses and irradiation. The measurement was obtained by monitoring of a pH meter after extracting a specific volume of solution at that time. The effect of H₂O₂ concentration on the Fenton-like process was analyzed over a wide range (1 mM–1 M) of UV-Vis irradiation using a 300 W xenon simulated solar light lamp (Perfectlight Scientific Pty Ltd., Beijing, China) applied to enhance the catalytic behavior of the Fe-based metallic glass ribbons. According to the specific absorbance peak (λ_{max}) of dyes in the visible spectrum, the decolorization of dyes could be quantified by UV-Vis spectrometry (Lambda 35, Perkin Elmer, Shelton, CT, USA), where each sample (3.5 mL) was extracted from reacting solution and characterized immediately at predetermined time intervals. Total organic carbon (TOC) removal rates before and after the Fenton-like process were measured by a TOC analyzer (TOC-LCPH, Shimadzu, Sydney, Australia).

4. Conclusions

In summary, a wide range of dyes including azo dyes (i.e., MO, BB, BY and BR), a thiazine dye (i.e., MB), triarylmethane dyes (i.e., MG and CV) and a nitroso dye (i.e., NG) have been selected to investigate their degradation by two Fe-based (Fe₇₈Si₉B₁₃ and Fe₇₃.₅Si₁₃.₅B₉Cu₁Nb₃) metallic glasses activated Fenton-like processes, which aims to get a systematic understanding of the correlation of the metallic glasses with dye degradation, focusing on the thermal catalytic behavior, H₂O₂ concentration at different pH values, decolorization and mineralization. The following points shed light on the significance in this work:

- The correlation of the thermal catalytic behavior between metallic glasses and dyes has been revealed, where Fe₇₈Si₉B₁₃ metallic glass ribbons with active catalytic ability have a lower activation energy (ΔE) than that of Fe₇₃.₅Si₁₃.₅B₉Cu₁Nb₃ in all dye degradations, and the ΔE of the metallic glass depends on the reaction environment (dyes).
- Under acidic conditions, a high H₂O₂ concentration is unfavorable to catalytic activity. However, under neutral conditions, a photo-enhanced Fenton-like process catalyzed by Fe-based metallic
glasses facilitates dye degradation at higher concentrations of H$_2$O$_2$ due to the decrease of pH and enhancement of irradiation.

- Fe$_{78}$Si$_{B3}$ metallic glass ribbons achieve an average TOC removal rate of 42.7% dye degradation in 15 min, which is higher than Fe$_{73.5}$Si$_{13.5}$B$_6$Cu$_{1}$Nb$_3$ (12.6%). The decolorization rate is closely related to the mineralization rate.

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